# T08D01 – (8.2) ****Properties of Acids and Bases Notes****

Name ……………………………………………………..

1. 8.2.1 Outline the characteristic properties of acids and bases in aqueous solution. (2)

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| **Properties of Acids** | **Properties of Bases** |
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* 1. What is an electrolyte?
  2. What happens to a base in water?

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|  | **What happens?** | **Relevant Equations:** | **In general:** |
| **Acid with Metal** |  |  |  |
| **Acid with Metal Carbonates** |  |  |  |
| **Acid with Bases** |  |  |  |
| **Acid with Metal Oxides** |  |  |  |
| **Acid with Metal Hydroxides** |  |  |  |

* 1. In general, provide a diagram (web) for the generic reactions acids are able to undergo:
  2. Use the example of hydrogen chloride (s) dissolved in water and the organic material methyl benzene to describe why water is important to the function of acids:

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| **Test** | **Solution of HCl in water** | **Solution of HCl in C6H5CH3** |
| **Universal Indicator Paper** |  |  |
| **Addition of CaCO3** |  |  |
| **Conductivity** |  |  |
| **Enthalpy of Solution** |  |  |

* 1. What is the difference between “hard” and “soft” water?
  2. What is the difference between a “normal salt” and an “acid salt?”